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Synthesis and Characterization of Co-doped Manganese Ferrite Nanoparticles

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ABSTRACT

$\text{Mn}_{1-x}\text{Co}_x\text{Fe}_2\text{O}_4$ ($x=0, 0.25, 0.5, 0.75,$ and 1) nanoparticles were synthesized using the sol-gel technique. The effect of Co substitution on the structural, elastic, and magnetic properties was systematically studied using X-ray diffraction (XRD) spectroscopy, Fourier transform infrared (FTIR) spectroscopy, and vibrating sample magnetometer (VSM). XRD patterns elucidate the formation of the cubic spinel phase for all prepared samples. Secondary Mn-oxid phases were observed for samples with low Co content $x<0.5$. All the samples fitted were with Rietveld refinement using Maud software to obtain the lattice parameter (a_{exp}) and particle size (D_{Maud}). Values of the lattice parameter were calculated theoretically and showed a good agreement with a_{exp} . The particle size was calculated from the Williamson and Hall plot and then compared to D_{Maud} , where considerable differences were observed. FTIR measurements confirm the formation of the cubic spinel phase and the elastic moduli were calculated. The magnetic parameters showed a strong dependence on the Co content. The saturation magnetization (M_s) and coercivity (H_c) increased by increasing x . By fitting the magnetic hysteresis loops of different $\text{Mn}_{1-x}\text{Co}_x\text{Fe}_2\text{O}_4$ nanoparticles, the effective magnetic anisotropy constant K and the anisotropy field (H_a) were obtained.

1. INTRODUCTION

Developing nanodevices, molecular apparatuses, and nano spectroscopies attract interest starting from the fabrication methods and characterization techniques to the final devices' applications. The ferrite nanoparticles have promising applications in the field of molecular transistors, molecular resonant tunneling diodes, memory, and configurational switch, and charged-based memory [1,2]. A typical spinel ferrite has the composition $(\text{M}_{1-\delta}\text{Fe}_\delta)_\text{A}[\text{M}_\delta\text{Fe}_{2-\delta}]_\text{B}\text{O}_4$, where A refers to the tetrahedral site and B is the octahedral site. M denotes divalent ions of the transition elements. The degree of inversion (δ) represents the concentration of divalent ions in the octahedral site. For a fully inverted spinel ferrite, all the divalent ions occupying the B-site and trivalent Fe^{3+} ions are equally distributed between the A- and B- sites [3-6]. M can be a mixture of divalent ions, whereas, transition or/and rare earth trivalent ions partially substitute Fe^{3+} ions [7-10]. Electric and magnetic properties of the ferrite compounds critically depend on the chemical composition, preparation conditions, particle size, and the distribution of

the cations between the tetrahedral, and octahedral sites [11-13]. MnFe_2O_4 ferrites are considered a partial inverse spinel structure, with an intense octahedral site preference of Mn^{2+} ions. On the other side, CoFe_2O_4 ferrites have a complete inverse spinel structure. Owing to the differences in the size, magnetic moments, and spin orbital coupling between Mn^{2+} and Co^{2+} ions, the magnetic properties of the two ferrites vary significantly. Cobalt ferrite is a hard magnetic material with a large coercive field whereas, Mn ferrite is soft and has a relatively low coercive field. Interestingly, doping is not the only factor that can alter the magnetic properties of ferrite compounds. Different preparation methods and heat treatments of the prepared samples yield particles of different sizes and particle shapes, which will consequently affect the magnetic properties [11,14,15].

In this work, $\text{Mn}_{1-x}\text{Co}_x\text{Fe}_2\text{O}_4$ nanoparticles were prepared using the Sol-gel technique. The influence of increasing Co content on the structural, elastic, and magnetic properties was systematically studied. Both Mn^{2+} and Co^{2+} ions are divalent; however, they have different ionic radii and magnetic moments.

Accordingly, the changes in the structural and magnetic properties are expected. The crystal structure was investigated using X-ray powder diffraction. The Fourier-transform infrared spectroscopy (FTIR) was used to investigate the elastic properties of the prepared samples. The characterization of the magnetic properties was achieved by performing vibrating sample magnetometer (VSM) measurements.

2 Experimental

2.1 Materials and methods

All the studied ferrites nanoparticles were synthesized using the sol-gel citrate precursor method as described in Figure (1). The nitrates of different elements; $\text{Fe}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$, $\text{Mn}(\text{NO}_3)_2 \cdot 4\text{H}_2\text{O}$, and $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$, citric acid ($\text{C}_6\text{H}_8\text{O}_7 \cdot \text{H}_2\text{O}$), and distilled water were used in the preparation procedures. The compounds are prepared by dissolving nitrates and citric

acid in an appropriate amount of distilled water. A complete dissolution of the mixture occurs through stirring onto a hot plate at $80\text{ }^\circ\text{C}$. The gels are heated on a hot plate until the complete evaporation of the aqueous solution and fine ashes are obtained. The dried powder was thermally treated at $400\text{ }^\circ\text{C}$ for 4 h.

The purity and structure of the prepared samples were investigated using X-ray diffraction (XRD) (Philips X'pert multipurpose diffractometer) with CuK_α radiation ($\lambda = 1.5418\text{ \AA}$). The crystal structure was refined and the Rietveld profile method was applied using the Maud software [16]. The Fourier transform infrared (FTIR) measurements of different $\text{Mn}_{1-x}\text{Co}_x\text{Fe}_2\text{O}_4$ samples were performed using Spectrum 100 (Perkin Elmer, USA) in the spectral range of $1000\text{--}250\text{ cm}^{-1}$. The magnetic properties were investigated using a vibrating sample magnetometer VSM (9600-1 LDJ, USA) at room temperature with an applied magnetic field up to 20 kG.

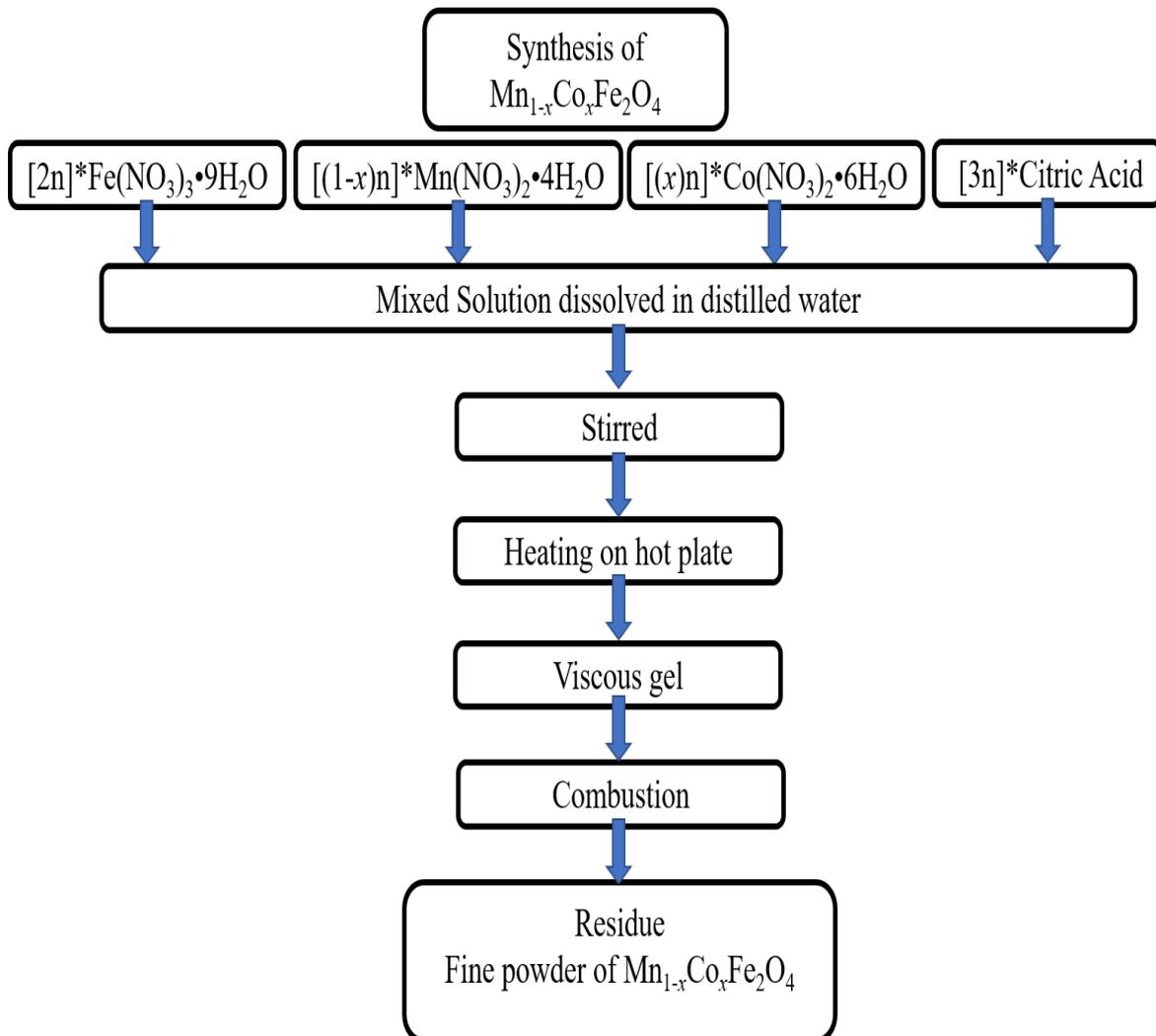


Fig. (1): Flowchart of synthesizing of $\text{Mn}_{1-x}\text{Co}_x\text{Fe}_2\text{O}_4$ nanoparticles

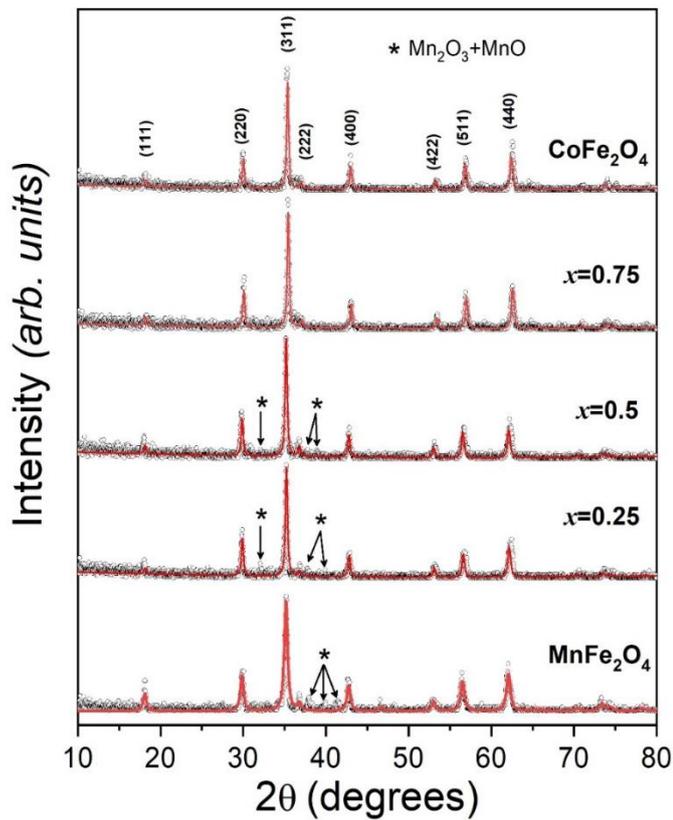


Fig. (2): XRD patterns of $Mn_{1-x}Co_xFe_2O_4$ ($x=0, 0.25, 0.5, 0.75$ and 1) nanoparticles

3. RESULTS AND DISCUSSION

3.1. XRD measurements

The XRD patterns of the $Mn_{1-x}Co_xFe_2O_4$ nanoparticles are shown in Figure (2). All the samples are crystallized in the cubic spinel ferrite phase with the $Fd-3m$ space group. For samples with a high Mn concentration, $x \leq 0.5$, additional diffraction peaks were observed and identified as Mn oxide phases (MnO and

Mn_2O_3). The intensity of the secondary phase peaks decreases with increasing the Co content, and these peaks disappeared in the $Mn_{0.25}Co_{0.75}Fe_2O_4$ and $CoFe_2O_4$ samples. The lattice parameter (a) obtained from the Rietveld analysis is listed in Table (1) and shown in Figure (3a). The lattice parameter slightly decreases by introducing Co^{2+} ions to the unit cell up to $x=0.5$, whereas for further increasing of the Co content, a sharply decreases. The particle size gradually increases by increasing x (Table 2 and Figure 3b). Using the cation distribution between the A- and B- sites, the theoretical lattice parameter a_{th} was calculated. The mean ionic radius per molecule of the A-site (r_A) and B-site (r_B) is given by [8]:

$$r_A = C_{Fe}r(Fe^{3+}) \quad (1)$$

$$r_B = \frac{1}{2}[C_{Fe}r(Fe^{3+}) + C_{TM}r(TM^{2+})] \quad (2)$$

Where; $r(Fe^{3+})$ is the ionic radius of Fe^{3+} in the tetrahedral and octahedral sites, 0.49 Å and 0.645 Å, respectively. The $r(TM^{2+})$ is the ionic radius of TM^{2+} (Co^{2+} , and Mn^{2+}) in the octahedral site, (0.745 Å and 0.83Å), respectively [17]. The concentration of Fe^{3+} and TM^{2+} ions in A- and B- sites are C_{Fe} and C_{TM} , respectively. The average (A - O), and (B - O) bond lengths are R_A and R_B , respectively and can be expressed as [8]:

$$R_A = r_A + r_O \quad (3)$$

$$R_B = r_B + r_O \quad (4)$$

Where $r_O = 1.32$ Å is the radius of oxygen ion. The lattice parameter a_{th} can be obtained from [8]:

$$a_{th} = \frac{8}{3\sqrt{3}}[(r_A + r_O) + \sqrt{3}(r_B + r_O)] \quad (5)$$

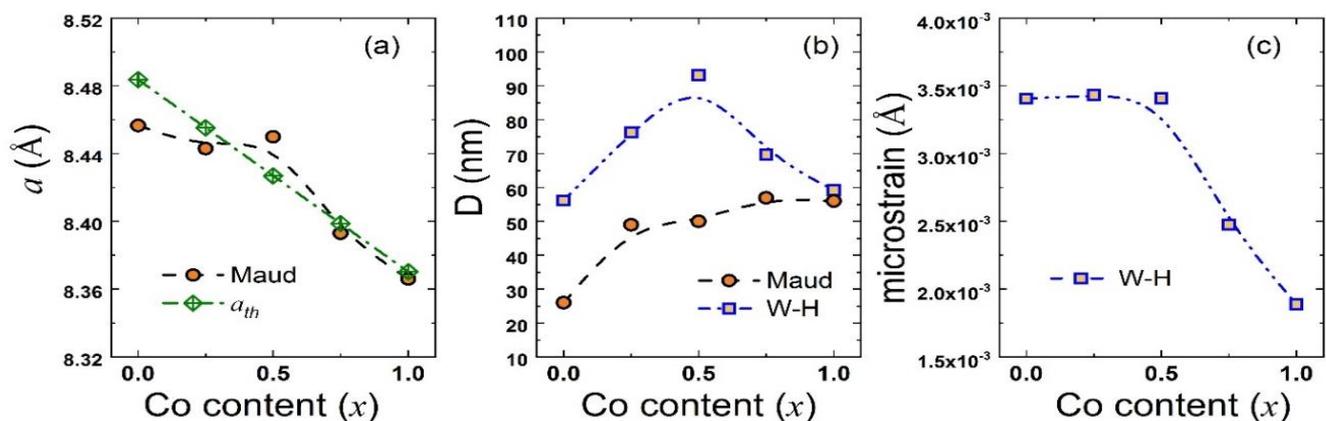


Fig. (3 a): The lattice parameter a , b) particle size D calculated from W-H plots, c) Microstrain ϵ of different $Mn_{1-x}Co_xFe_2O_4$ nanoparticles

The values of the structure parameters; a_{th} , R_A , and R_B , are listed in Table (1). The ionic radius of Co^{2+} ion is smaller than that of Mn^{2+} ion, and therefore, as shown in Figure (3a), a_{th} decreases linearly with increasing the Co content. The values of the theoretically calculated lattice parameter (a_{th}) are close to the values of a_{exp} . The trend of the decrease of both a_{th} and a_{exp} is similar for $x > 0.5$. The differences observed in the trend for $x < 0.5$ can be attributed to the existence of the secondary phase, which might cause an interstitial substitution and increase the uncertainty in Maud fitting. The decrease of the lattice parameter by

increasing the Co content can be attributed to the differences in the ionic radii of the Mn^{2+} and Co^{2+} ions, whereas, $r(Co^{2+}) < r(Mn^{2+})$. The X-ray density is given by:

$$\rho_x = \frac{8M}{N_A V} \quad (6)$$

Where M is the molecular weight and $V = a^3$ is the volume of the unit cell. The lattice parameter decreases with increasing x and therefore the density increases gradually with the increase of the Co content.

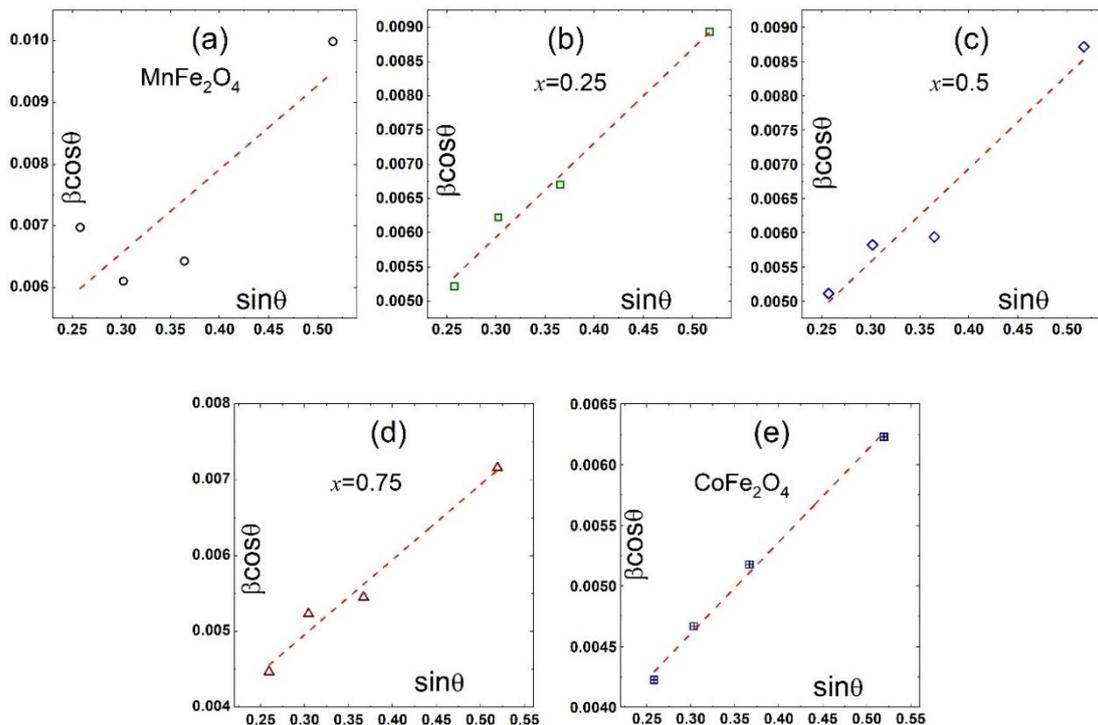


Fig. (4): W-H plots of different $Mn_{1-x}Co_xFe_2O_4$ nanoparticles

Table (1): Average bond lengths (R_A , and R_B), experimental (a_{exp}), theoretical (a_{th}) lattice constants, density (ρ_x), the hopping lengths (L_A , and L_B), average particle size obtained from Maud fitting (D_{Maud}), and W-H plots (D_{W-H}), the microstrain (ϵ) of $Mn_{1-x}Co_xFe_2O_4$ nanoparticles

x	R_A (Å)	R_B (Å)	a_{th} (Å)	a_{exp} (Å)	ρ_x (gcm ⁻¹)	L_A (Å)	L_B (Å)	D_{Maud} nm	D_{W-H} nm	ϵ (10 ⁻⁴) (Å)
0	1.86	2.108	8.483	8.457	5.07E3	3.662	2.99	26	56	34.05
0.25	1.86	2.097	8.455	8.443	5.11E3	3.656	2.985	49	76	34.325
0.5	1.86	2.086	8.427	8.45	5.12E3	3.659	2.988	50	93	34.075
0.75	1.86	2.076	8.399	8.393	5.25E3	3.634	2.967	57	69	24.75
1	1.86	2.065	8.37	8.366	5.32E33	3.623	2.958	56	59	18.875

The average crystallite size and microstrain ‘ ε ’ values can be obtained from Williamson and Hall (W-H) equation [18]:

$$\beta \cos\theta = \frac{\lambda k}{D} + 4\varepsilon \sin\theta \quad (7)$$

Where β is the broadening of the XRD peaks, λ is the XRD wavelength (CuK α radiation with $\lambda = 1.5418$), $k=0.9$, ε microstrain, and D is the particle size. For all the Mn $_{1-x}$ Co $_x$ Fe $_2$ O $_4$ ($x=0, 0.25, 0.5, 0.75,$ and 1) nanoparticles, XRD peaks (2 2 0), (3 1 1), (2 2 2), (4 0 0) and (4 4 0) were fitted using Gaussian line profiles and from the fitting, values of θ and β were obtained. Figure (4) shows Williamson and Hall plot for all samples. The values of the particle size obtained from the W-H plot are listed in Table (1) and compared to that obtained from Maud fitting in Fig. (3b). A considerable difference was observed in both the values and trends of D_{Maud} and $D_{\text{W-H}}$. This difference can be related to the observation of the secondary Mn oxide phases for $x<0.5$, which influences the width of the XRD peaks and complicate the calculation of the particle size. The values of the microstrain have a behavior similar to the lattice parameter (Fig. 3c), whereas ε shows a clear decrease for $x>0.5$. The decrease of the microstrain can be considered as a result of substituting Mn $^{2+}$ ions with smaller Co $^{2+}$ ions [19]. Simultaneously, the coexistence of multi-crystallographic phases causes additional strain. This strain is relieved for single phases Mn $_{0.25}$ Co $_{0.75}$ Fe $_2$ O $_4$ and CoFe $_2$ O $_4$ nanoparticles.

The hopping lengths, L_A , and L_B represent the distance between the magnetic ions in the tetrahedral sites and octahedral sites, respectively. L_A and L_B can be calculated as [20]:

$$L_A = \frac{\sqrt{3}}{4} a \quad (8)$$

$$L_B = \frac{\sqrt{2}}{4} a \quad (9)$$

The values L_A and L_B are shown in Table (1). Because of the decreased lattice parameter, both lengths decrease for $x>0.5$. Such a decrease of L_A and L_B would logically strengthen the magnetic interaction with increasing the Co content.

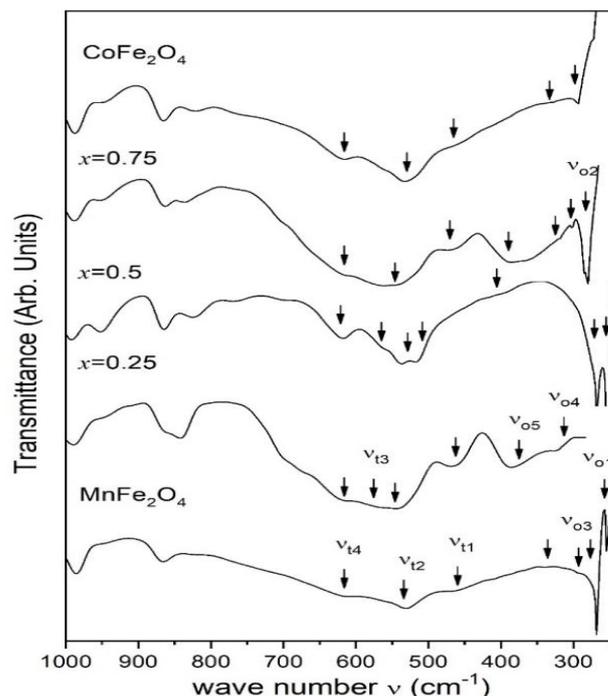


Fig (5): FTIR measurements of different Mn $_{1-x}$ Co $_x$ Fe $_2$ O $_4$ nanoparticles

3.2 Fourier-transform infrared (FTIR) spectroscopy

The FTIR spectra of the different Mn $_{1-x}$ Co $_x$ Fe $_2$ O $_4$ ($x=0, 0.25, 0.5, 0.75,$ and 1) nanoparticles are shown Figure (5). The two characteristic vibrational bands; ν_t (450–650 cm $^{-1}$) and ν_o (250–400 cm $^{-1}$) belong to the tetrahedral and octahedral sites, respectively. Each band contains more than one vibrational frequency, which can be attributed to the difference in the atomic masses of different ions (Fe, Mn, and Co), bond lengths, and the correlation between A-site complexes and the B-site [21–24]. The presence of secondary Mn oxide phases for samples Mn $_{1-x}$ Co $_x$ Fe $_2$ O $_4$ ($x=0, 0.25, 0.5$) will logically complicate the FTIR response and cause additional vibrational frequencies. The bond length in the A-site is shorter than that in the B-site, hence, stretching the cation-O bond in the A-site, which will require a higher energy than that required for the cation-O bond in the B-site [24,25]. This fact causes a higher tetrahedral vibrational band than that for the B band. The vibrational frequencies of all samples are listed in Table (2). The tetrahedral and octahedral sites force constants (k_t and k_o , respectively) can be calculated as [26]:

$$k_t = 7.62 \times M_A \times \nu_t^2 \times 10^{-7} \text{ N/m} \quad (10)$$

$$k_o = 10.62 \times \frac{M_A}{2} \times \nu_o^2 \times 10^{-7} \text{ N/m} \quad (11)$$

The average molecular weights of the cations in the tetrahedral and octahedral sites are M_A , and M_B and were calculated using the cation distribution

(Fe)_A[Mn_{1-x}Co_xFe]_BO₄. Figure (6a) illustrates the average force constant k_{tav} and k_{oav} in both A- and B-sites. Because of the shorter bond in the A-site, the bond in the tetrahedral site is stronger than that in the octahedral site and hence, $k_t > k_o$. The force constant is sensitive to the bond length, charge, mass, and size of the ions forming these bonds [25,27]. In the A-site, k_t changes with increasing x , however, no specific trend was observed. Alternatively, k_o increases slightly with increasing x up to 0.5 and then it slowly decreases.

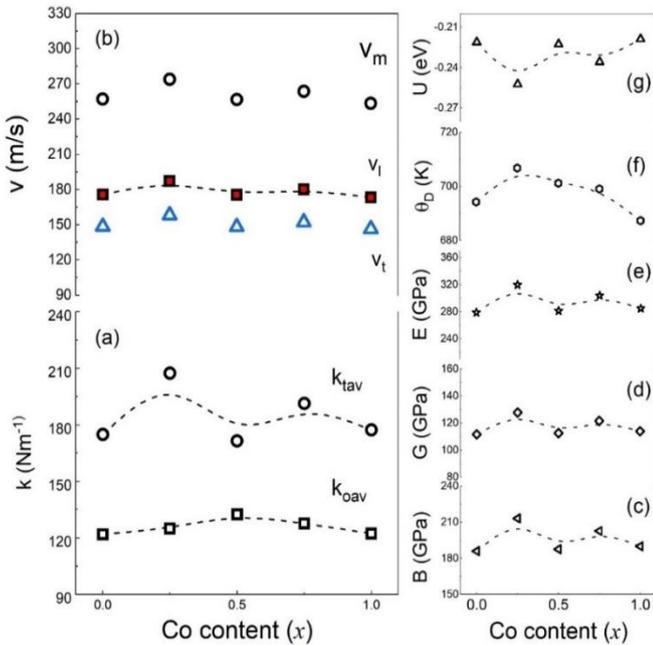


Fig. (6 a): Force constant k , b) velocity of elastic waves (V_l , V_t and V_m), c) Bulk modulus (B), d) Young's modulus (E) e) rigidity modulus (G). f) Debye temperature (θ_D) and g) lattice energy (U) of different elastic parameters of different $Mn_{1-x}Co_xFe_2O_4$ nanoparticles

The average stiffness relates to the bulk modulus (B) as:

$$B = \frac{k}{a} \quad (12)$$

Where a is the lattice constant obtained from the XRD measurements. Different elastic parameters can be calculated as [26-28]:

The longitudinal velocity of elastic waves (V_l)

$$v_l = \left[\frac{9B}{5\rho_x} \right]^{\frac{1}{2}} \quad (13)$$

The transverse velocity of elastic waves (V_t)

$$v_t = \left[\frac{3B}{5\rho_x} \right]^{\frac{1}{2}} \quad (14)$$

The mean velocity of elastic waves (V_m)

$$v_m = \left[\frac{1}{3} \left(\frac{2}{v_t^3} + \frac{1}{v_l^3} \right) \right]^{-\frac{1}{3}} \quad (15)$$

The rigidity modulus (G)

$$G = \rho_x v_t^2 \quad (16)$$

Young's modulus (E)

$$E = \frac{9BG}{3B + G} \quad (17)$$

The lattice energy (U)

$$U = -3.108 (Mv_m) \times 10^{-12} \quad (18)$$

Table (2): Tetrahedral and octahedral vibrational frequencies (ν_{t1} , ν_{t2} , ν_{tav} , ν_{o1} , ν_{o2} , ν_{o3} , and ν_{oav}) of the $Mn_{1-x}Co_xFe_2O_4$ nanoparticles

x	ν_{A1}	ν_{A2}	ν_{A3}	ν_{A4}	ν_{Aav}	ν_{B1}	ν_{B2}	ν_{B3}	ν_{B4}	ν_{B5}	ν_{Bav}
	(cm ⁻¹)										
0	462	-	530.4	613.3	535.3	254.6	265.7	293.2	341.4	387.7	308.5
0.25	468.7	-	542	615.3	542	-	-	297	324.9	385.8	335.9
0.5	-	516.9	538.1	619.1	558	254.6	268.1	-	-	393.5	305.4
0.75	470.6	-	555.5	617.2	547.8	279.7	-	302.8	322.1	385.8	322.6
1	461	-	532.3	615.3	536.2	293.2	-	-	327.9	-	310.5

The values of the longitudinal (V_l), transverse (V_t), mean (V_m) velocity, and the elastic parameters; B, G, E, and U are shown in Figure (6). Three main factors that govern the behavior of all these parameters are as follows:

- The change in the lattice parameter will affect the density and length of the bonds in both sites, as shown before in the XRD measurements. As motioned before, a decreases with increasing the Co content, however, this decrease is not uniform. The decreasing of the lattice parameter tends to increase the force constant and consequently, the elastic parameters; B, G, and E.
- The particle size D is higher for the Co substituted samples. Increasing the particle size tends to decrease the elastic moduli; B, G, and E [21].
- The presence of the secondary Mn oxide phases for samples $Mn_{1-x}Co_xFe_2O_4$ ($x=0, 0.25, 0.5$) will affect the elastic moduli.

For all the calculated parameters, V_l , V_t , V_m , B, G, and E, no specific dependence on x was observed. Such behavior suggests that the motioned i,ii and iii factors are variable for each sample, non-uniform and the resultant effect has no systematic dependence on the Co content.

The Debye temperature (θ_D) can be obtained from the Waldron method as [26]:

$$\theta_D = \frac{hcv_{av}}{k} \quad (19)$$

Where h is the Planck's constant, c is the velocity of light, and k_B is the Boltzmann constant. θ_D increases with the initial doping of Co^{2+} ions ($x=0.25$) and then it decreases gradually. As observed before for $CoFeAlO_4$, Debye decreases with the increase of the particle size, whereas, smaller ferrite nanoparticles have stiffer crystals [21].

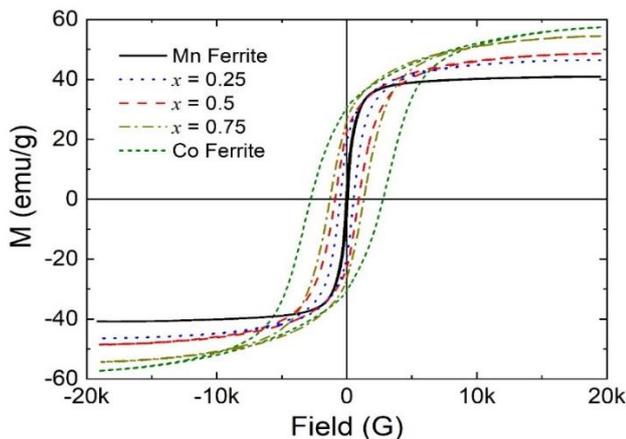


Fig (7): Room temperature hysteresis loops of different $Mn_{1-x}Co_xFe_2O_4$ nanoparticles

3.3 Vibrating sample magnetometer (VSM) measurements

Room temperature hysteresis loops of different $Mn_{1-x}Co_xFe_2O_4$ nanoparticles are shown in Figure (7).

The saturation magnetization (M_s), coercivity (H_c), remanent magnetization (M_r), and the squareness ratio are listed in Table (3). The saturation magnetization increases linearly with increasing x . M_s can be expressed in the Bohr magneton unit as:

$$M_{obs} = \frac{(\text{saturation magnetization}) \times (\text{molecular weight})}{5585} \quad (20)$$

The values of the observed moment M_{obs} are shown in Table (3) and Figure (8). The magnetic properties of the ferrites originated mainly by A-A, A-B, and B-B magnetic super exchange interactions. According to Néel's Model, the magnetic moment M_{cal} is described by[8]:

$$M_{cal} = M_B - M_A$$

The tetrahedral moment M_A and octahedral moment M_B were calculated using the cation distribution $(Fe)_A[Mn_{1-x}Co_xFe]_B O_4$ and the magnetic moment of Co^{2+} , Fe^{3+} ions are $2 \mu_B$ and $5 \mu_B$, respectively. Mn^{2+} ions with (d^5) can have two magnetic moments; high spin with $5 \mu_B$ and low spin with $1 \mu_B$. The low values M_{obs} indicate the low spin state of Mn^{2+} ions. The values of M_{cal} are presented in Table (3) and Figure (8). M_r also increases gradually with the increase in x , but with a faster rate than that of M_s . Therefore, the squareness ratio M_r/M_s increases gradually with an increase of the Co content and reflects a gradual increase of the particle size. The values of the squareness suggest that different $Mn_{1-x}Co_xFe_2O_4$ nanoparticles have uniaxial anisotropy, which arises due to surface effects [29].

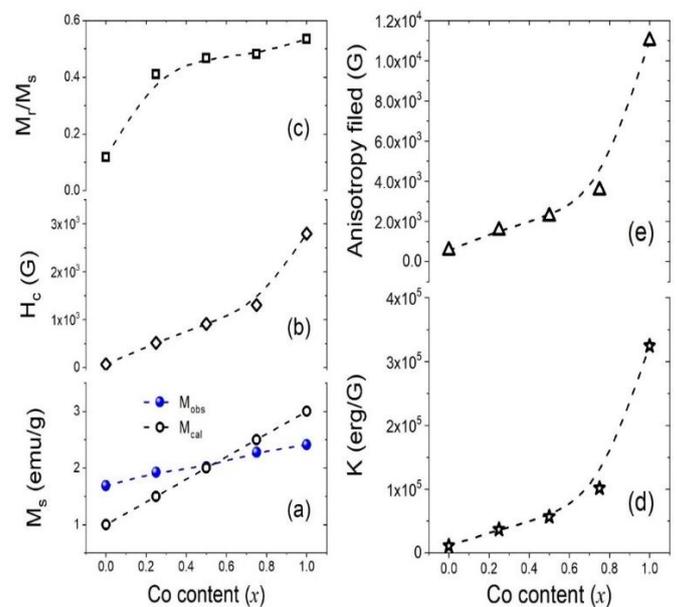


Fig. (8): Magnetic parameters of different $Mn_{1-x}Co_xFe_2O_4$ nanoparticles a) The observed moment (M_{obs}), and calculated moment (M_{cal}), b) Coercivity (H_c), c) Squareness ratio, d) Effective magnetic anisotropy constant (K), and e) Anisotropy field (H_a)

The coercivity (H_c) increases dramatically with increasing x , from 66 G for $MnFe_2O_4$ to almost 2794 G for $CoFe_2O_4$. Such an increase in coercivity suggests a gradual move from soft ferrite to hard ferrite by increasing the Co content. $M(H)$ can be fitted by the law approach to saturation as [30]:

$$M(H) = M_s \left(1 - \frac{a}{H} - \frac{b}{H^2} \dots \right) + \chi_0 H \quad (21)$$

For a uniaxial symmetry, the coefficient b relates to the effective magnetic anisotropy constant [30]:

$$b = \frac{4}{15} \left(\frac{K}{\mu_0 M_s} \right)^2 \quad (22)$$

The anisotropy field (H_a) can be obtained from [30]:

$$H_a = \frac{2K}{\mu_0 M_s} \quad (23)$$

The fitting of hysteresis loops of different $Mn_{1-x}Co_xFe_2O_4$ nanoparticles are shown in Figure (9). The calculated effective magnetic anisotropy constant (K) is presented in Table (3) and Figure (8d). The strong increase of H_c is mainly originated by the enhanced values of K and H_a with increasing the Co content. Both K and H_a depend on the type and concentration of magnetic ions in the ferrite system. Even more, the particle size and shape strongly affect the values of the magnetic anisotropy constant and the anisotropy field. All these factors are present in the $Mn_{1-x}Co_xFe_2O_4$ nanoparticles. Accordingly, the changes in the values of the magnetic parameters can be attributed to the interplay between the change in the Co content and the particle size.

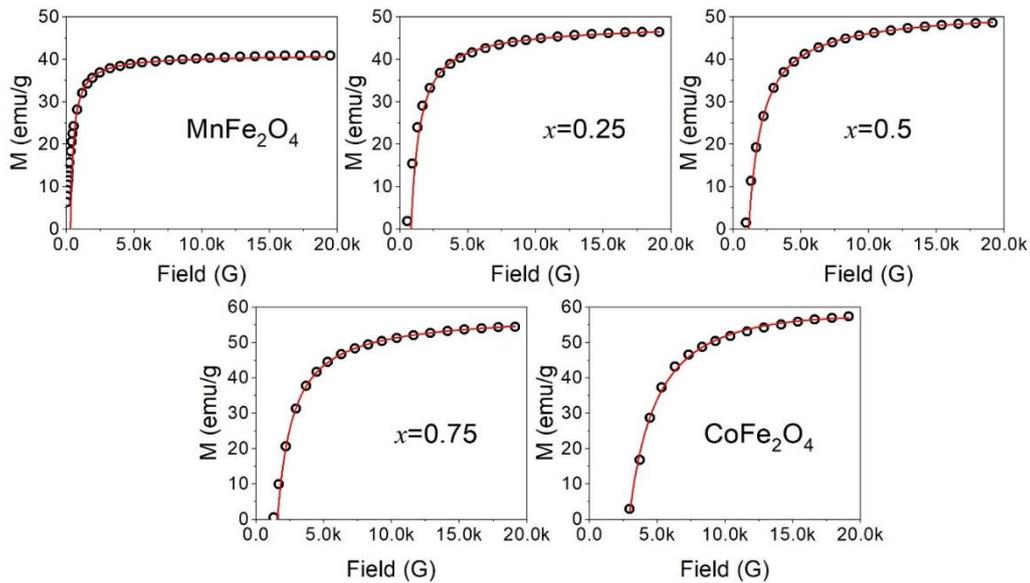


Fig. (9): Fitting of room temperature hysteresis loops of different $Mn_{1-x}Co_xFe_2O_4$ nanoparticles using Eq. (21)

Table (3): saturation magnetization (M_s), coercivity (H_c), remanent magnetization (M_r), the squareness ratio ($R^2=M_r/M_s$), observed moment M_{obs} , calculated moment M_{cal} , fitting parameters $M_s(fit)$, and anisotropy constant (K), and anisotropy field (H_a) of $Mn_{1-x}Co_xFe_2O_4$ nanoparticles

		0	0.25	0.5	0.75	1
M_s	(emu/g)	40.88	46.46	48.59	54.43	57.38
H_c	(G)	66.36	515.88	906.04	1302.7	2793.5
M_r	(emu/g)	4.84	19.05	22.68	26.22	30.66
M_r/M_s		0.118	0.41	0.467	0.482	0.534
M_{obs}	(μB)	1.69	1.93	2.02	2.28	2.41
M_{cal}	(μB)	1	1.5	2	2.5	3
M_s (Fit)	(emu/g)	40.87	47.68	50.41	57.29	58.99
K	(erg/G)	11288.3	37078.6	56990.8	101770.1	324425.1
H_a	(G)	552.37	1555.3	2261.3	3552.73	11000.3

4. CONCLUSION

Mn_{1-x}Co_xFe₂O₄ nanoparticles were successfully synthesized using the Sol-gel technique. The effect of increasing the Co content on the structural, elastic and magnetic properties were investigated. The X-ray powder diffraction measurements confirmed the formation of the face center cubic spinel phase. The secondary phase in the samples ($x \leq 0.5$) disappeared with increasing the cobalt content. The experimental lattice parameter a_{exp} , obtained by fitting the XRD patterns using Maud program, decreased by increasing x , and showed a good agreement with the theoretically calculated a_{th} . The differences in the values and behavior of the particle size obtained from W-H plots (D_{W-H}) and the particle size obtained from the Maud fitting can be attributed to the existence of the secondary Mn oxide phases for $x < 0.5$. FTIR spectra revealed the formation of the cubic spinel phase, whereas, the tetrahedral and octahedral vibrational bands were observed. The magnetic parameters; saturation magnetization (M_s), coercivity (H_c), remanent magnetization (M_r) and the squareness ratio increased with increasing the Co content. The values of M_s suggested that Mn²⁺ ions are in a low spin state ($1 \mu_B$), therefore, increasing x will logically increase the saturation magnetization. The strong increase in the coercivity values with increasing x confirmed the transition from soft Mn-ferrite toward hard Co-doped Mn ferrite system.

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